

3.4 Green

3.4.1 Chromium Oxide Green, Modified Chromium Oxide Green

Chromium oxide green Cr_2O_3 (**G01** - **G02**) exhibits strong scattering alternating with strong absorption across the visible spectrum, and strong scattering and mild absorption in the NIR. Since the pigment is almost opaque in the visible, a thin layer of chromium oxide green over a white background yields a medium-green coating with good NIR reflectance (0.57 for 13- μm thick film G02). The modified chromium oxide green (**G03**) is mostly chromium oxide, with small amounts of iron oxide, titanium dioxide, and aluminum oxide [16]. A layer of the modified chromium oxide green over a white background produces a medium green with excellent NIR reflectance (0.71).

Cr_2O_3 green is often mentioned as an infrared-reflective pigment that is useful for simulating the high infrared reflectance of plant leaves. Indeed, a high NIR reflectance is observed. However, our data for sample films G01 and G02 do show that there is a broadband absorption of about 10 mm^{-1} in the near-infrared. While our measurements of absorptance coefficient are not precise for low absorptances, this value is clearly distinct from zero. Pure Cr_2O_3 , fired in air, tends to become slightly rich in oxygen, which results in p-type semiconducting behavior [23, 24]. Thus it is possible that the broadband IR absorption of Cr_2O_3 is due to free carrier absorption by mobile holes. Ref. [23] also reports that doping with Al can reduce the p-type conductivity in Cr_2O_3 , so it seems likely that doping with Al and/or certain other metals can also reduce the IR absorption.

The modified chromium oxide green G03 is similar to G01 and G02 Cr_2O_3 . However its green reflectance peak at 550 nm is somewhat smaller and its infrared absorption is clearly much smaller than those of samples G01 and G02.

3.4.2 Cobalt Chromite Green

Cobalt chromite green (**G04** - **G06**) is similar to cobalt chromite blue, and is commonly used for military camouflage.

3.4.3 Cobalt Titanate Green

Cobalt titanate green (**G07** - **G09**) is similar to cobalt chromite green, but scatters more strongly across the entire solar spectrum and has a pronounced absorption trough around 500 nm. A white background makes cobalt teal G07 very NIR reflective (0.73) but also appear light blue (hence, the name teal). The other two cobalt titanate greens (G08, G09) have respectable NIR reflectances (0.47, 0.37) over white and appear medium green.

3.4.4 Phthalocyanine Green

Phthalocyanine green (**G10** - **G11**) is similar to phthalocyanine blue, but absorbs more strongly in the short NIR. Hence, the NIR reflectance of a thin phthalo green film over white, while respectable, is only 70% of that achieved by a thin layer of phthalo blue over white (0.45 for G10 vs. 0.63 for U12). Note also that the error in predicted reflectance over white for G11 is large, as discussed in the companion article [1].

References

- [1] Ronnen Levinson, Paul Berdahl, and Hashem Akbari. Solar spectral optical properties of pigments, Part I: model for deriving scattering and absorption coefficients from transmittance and reflectance measurements. *Solar Energy Materials & Solar Cells (accepted)*, 2004.
- [2] Peter A. Lewis. *Pigment Handbook*, volume I. John Wiley and Sons, 1988.
- [3] Gunter Buxbaum. *Industrial Inorganic Pigments*. Wiley-VCH, 2nd edition, 1998.
- [4] Willy Herbst and Klaus Hunger. *Industrial Organic Pigments*. VCH, 1993.
- [5] Y.S. Touloukian, D.P. DeWitt, and R.S. Harnicz. *Thermal Radiative Properties: Coatings*, volume 9 of *Thermophysical Properties of Matter*. IFI/Plenum, 1972.
- [6] Ralph Mayer. *The Artist's Handbook of Materials and Techniques*. Viking Penguin, 5th edition, 1991.
- [7] Society of Dyers and Colourists and American Association of Textile Chemists and Colorists. Colour index international: Fourth online edition. <http://www.colour-index.org>.
- [8] B. R. Palmer, P. Stamatakis, C. G. Bohren, and G. C. Salzman. A multiple-scattering model for opacifying particles in polymer films. *Journal of Coatings Technology*, 61(779):41–47, 1989.
- [9] E.S. Thiele and R.H. French. Computation of light scattering by anisotropic spheres of rutile titania. *Adv. Mater.*, 10(15):1271–1276, 1998.
- [10] Paul Berdahl. Pigments to reflect the infrared radiation from fire. *Journal of Heat Transfer*, 117:355–358, May 1995.
- [11] D.J. Rutherford and L.A. Simpson. Use of a flocculation gradient monitor for quantifying titanium dioxide pigment dispersion in dry and wet paint films. *Journal of Coatings Technology*, 57(724):75–84, May 1985.
- [12] R.R. Blakey and J.E. Hall. *Pigment Handbook*, volume I, chapter A (“Titanium Dioxide”), pages 1–42. John Wiley and Sons, 1988.
- [13] Daniel Russell Swiler. Manganese vanadium oxide pigments. U.S. Patent 6,485,557 B1, Nov 26 2002.
- [14] Dry Color Manufacturer's Association (DCMA). *Classification and chemical description of the complex inorganic color pigments*. Dry Color Manufacturer's Association, P.O. Box 20839, Alexandria, VA 22320, 1991.
- [15] E. B. Rabinovitch and J. W. Summers. Infrared reflecting vinyl polymer compositions. U.S. Patent 4,424,292, 1984.
- [16] Terrence R. Sliwinski, Richard A. Pipoly, and Robert P. Blonski. Infrared reflective color pigment. U.S. Patent 6,174,360 B1, Jan 16 2001.
- [17] V.A.M. Brabers. The electrical conduction of titanomagnetites. *Physica B*, 205:143–152, 1995.
- [18] L.B. Glebov and E.N. Boulos. Absorption of iron and water in the $\text{Na}_2\text{O-CaO-MgO-SiO}_2$ glasses, II. Selection of intrinsic, ferric, and ferrous spectra in the visible and UV regions. *J. Non-Crystalline Solids*, 242:49–62, 1998.

- [19] R.N. Clark. *Manual of Remote Sensing*, volume 3 (“Remote sensing for the earth sciences”), chapter 1 (“Spectroscopy of Rocks and Minerals, and Principles of Spectroscopy”), pages 3–58. John Wiley and Sons, <http://speclab.cr.usgs.gov>, 1999. Fig. 5.
- [20] R.J.H. Clark and D.G. Cobbold. Characterization of sulfur radical anions in solutions of alkali polysulfides in dimethylformamide and hexamethylphosphoramide and in the solid state in ultramarine blue, green, and red. *Inorganic Chemistry*, 17:3169–3174, 1978.
- [21] N.B. Mckeown. *Phthalocyanine Materials: Synthesis, Structure and Function*. Cambridge Univ. Press, Cambridge, UK, 1998.
- [22] S. Wilbrandt O. Stenzel A. Stendal, U. Beckers and C. von Borczyskowski. The linear optical constants of thin phthalocyanine and fullerite films from the near infrared to the UV spectral regions: estimation of electronic oscillator strength values. *J. Phys. B*, 29:2589–2595, 1996.
- [23] D. de Cogan and G.A. Lonergan. Electrical conduction in Fe_2O_3 and Cr_2O_3 . *Solid State Communications*, 15:1517–1519, 1974.
- [24] Hamnett Goodenough. *Landolt-Bornstein Numerical Data and Functional Relationships in Science and Technology, New Series, Group III: Crystal and Solid-State Physics*, volume 17g (Semiconductors: Physics of Non-Tetrahedrally Bonded Binary Compounds III), chapter 9.15.2.5.1: Oxides of chromium, pages 242–247,548–551. Springer-Verlag, Berlin, 1984.
- [25] G.B. Smith, A. Gentle, P. Swift, A. Earp, and N. Mronga. Coloured paints based on coated flakes of metal as the pigment, for enhanced solar reflectance and cooler interiors: description and theory. *Solar Energy Materials & Solar Cells*, 79(2):163–177, 2003.
- [26] G.B. Smith, A. Gentle, P. Swift, A. Earp, and N. Mronga. Coloured paints based on iron oxide and solicon oxide coated flakes of aluminium as the pigment, for energy efficient paint: optical and thermal experiments. *Solar Energy Materials & Solar Cells*, 79(2):179–197, 2003.